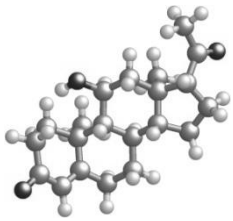

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Class: ……

Date: …… **/** …… **/** ……

Multiple-choice Questions on Acids, Bases and Salts #6

1. Which mixture would react with dilute sulfuric acid to form two different gases?

A Copper and magnesium carbonate

B Copper(II) carbonate and magnesium

C Copper(II) carbonate and magnesium oxide

D Copper(II) oxide and magnesium

2. Which salts are soluble in water?

1 ammonium carbonate, $(NH_4)_2CO_3$

2 calcium carbonate, $CaCO_3$

3 lead(II) carbonate, $PbCO_3$

4 sodium carbonate, Na_2CO_3

A 1 only

B 1 and 2 only

C 1 and 4 only

D 2 and 3 only

3. Which compound in a 1 mol/dm^3 solution has the lowest pH value?

A Ethanoic acid

B Hydrogen chloride

C Sulfuric acid

D Sodium hydroxide

4. Salts containing which of the following anions are always soluble in water?

A Carbonates

B Chlorides

C Nitrates

D Sulfates

5. What is a property of the hydroxide, OH^-, ion?

A It combines with hydrogen to form water.

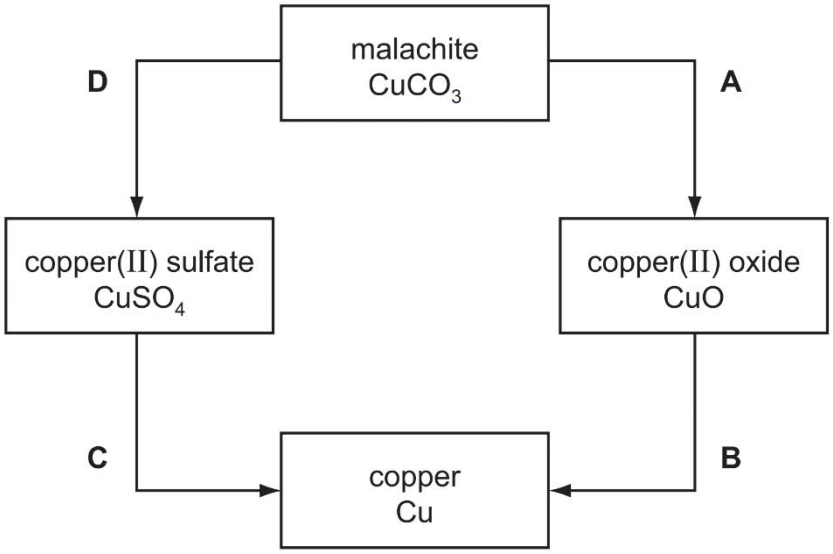
B It is present in water.

C It readily breaks down into hydrogen ions and oxide ions.

D It travels to the cathode in electrolysis of an aqueous solution.

6. The diagram shows some reactions of copper compounds.

Which change is made by adding an acid?



7. When the product of a reaction between two gases is added to water, a solution of pH 7 is formed. Which could be these gases?

A Hydrogen and chlorine

B Hydrogen and nitrogen

C Hydrogen and oxygen

D Oxygen and carbon monoxide

8. Which reagent is added to aqueous potassium chloride to prepare lead(II) chloride?

A Aqueous lead(II) nitrate **B** Lead

C Lead(II) carbonate **D** Lead(II) sulfate

9. Which statement about amphoteric oxides is **not** correct?

A They dissolve in water.

B They are formed only by metals.

C They react with aqueous sodium hydroxide to give salts.

D They react with aqueous acids to give salts.

10. To which substance is dilute sulfuric acid added to prepare lead(II) sulfate?

A Aqueous lead(II) nitrate

B Lead foil

C Powdered lead(II) carbonate

D Powdered lead(II) oxide

11. Which pair of substances reacts to form a salt and water only?

A Aqueous sodium chloride and aqueous silver nitrate.

B Aqueous sodium hydroxide and dilute ethanoic acid.

C Aqueous sodium carbonate and dilute sulfuric acid.

D Zinc and dilute hydrochloric acid.

12. Which reaction does **not** involve neutralisation?

A $H_2SO_4(aq) + 2NH_3(aq) \rightarrow (NH_4)_2SO_4(aq)$

B $H_2SO_4(aq) + BaCl_2(aq) \rightarrow BaSO_4(s) + 2HCl(aq)$

C $H_2SO_4(aq) + CuO(s) \rightarrow CuSO_4(aq) + H_2O(l)$

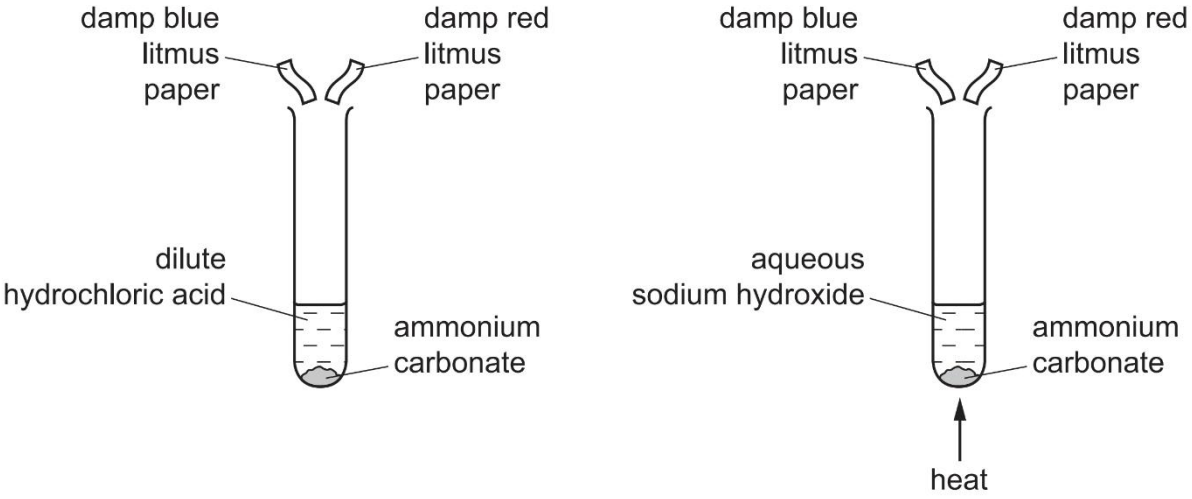
D $H_2SO_4(aq) + 2NaOH(aq) \rightarrow Na_2SO_4(aq) + 2H_2O(l)$

13. Two experiments were carried out.

In experiment 1, ammonium carbonate was reacted with dilute hydrochloric acid.

In experiment 2, ammonium carbonate was heated with aqueous sodium hydroxide.

In each experiment, the gas evolved was tested with damp blue litmus paper and damp red litmus paper.



experiment 1 experiment 2

Which row correctly shows the colour of both the pieces of litmus paper at the end of each experiment?

	experiment 1	experiment 2
A	blue	blue
B	blue	red
C	red	blue
D	red	red

14. The table shows some properties of four metal chlorides.

Which row is magnesium chloride?

	colour	solubility in water	method of preparation
A	green	insoluble	precipitation
B	green	soluble	metal and acid
C	white	insoluble	precipitation
D	white	soluble	metal and acid

15. Which row shows the pH values for 0.1 mol/dm^3 solutions of ammonia, hydrochloric acid, sodium chloride and sodium hydroxide?

	pH values			
	NH_3	HC*l*	NaC*l*	NaOH
A	1	7	13	11
B	7	1	11	13
C	11	1	7	13
D	13	11	7	1

16. A student has five reagents.

- dilute hydrochloric acid
- dilute sulfuric acid
- dilute nitric acid
- solid calcium carbonate
- solid copper(II) carbonate

How many soluble salts can be prepared?

A 3 **B** 4 **C** 5 **D** 6

17. Which statement is correct?

A Ammonia is produced when an ammonium salt is warmed with a dilute acid.

B Amphoteric oxides are oxides of certain metals.

C A neutral solution does not contain hydroxide ions.

D Soil with a high pH can be neutralised by adding lime, $Ca(OH)_2$.

18. Which reagent can be used to react with dilute hydrochloric acid to prepare silver chloride?

A Aqueous silver nitrate

B Solid silver

C Solid silver carbonate

D Solid silver oxide

19. Three separate mixtures of a solution and a solid are made, as shown in the table.

The mixtures are warmed.

In which mixtures does a gas form?

	NaOH(aq) and NH_4Cl(s)	H_2SO_4(aq) and NH_4Cl(s)	H_2SO_4(aq) and Mg(s)
A	✓	✓	✗
B	✓	✗	✓
C	✗	✓	✗
D	✗	✗	✓

key:

✓ = gas forms

✗ = no gas forms

20. Insoluble salts are prepared by reacting aqueous solutions of soluble salts. A precipitate forms. Which pairs of aqueous solutions form a precipitate?

1 barium chloride and nitric acid

2 barium chloride and sulfuric acid

3 barium nitrate and nitric acid

4 barium nitrate and sulfuric acid

A 1 and 2 only **B** 1 and 3 only

C 2 and 4 only **D** 3 and 4 only

21. What is a use of sulfuric acid?

A As a bleach.

B As a food preservative.

C In the manufacture of detergents.

D In the manufacture of vanadium(V) oxide, V_2O_5.

22. Which compound can be formed by precipitation?

A $AgCl$ **B** $CaCl_2$ **C** K_2SO_4 **D** $Pb(NO_3)_2$

23. Sodium hydroxide is added to a solution to alter its pH. A neutral solution is formed.

Which statement is correct?

A Sodium hydroxide is an acid and reacts with an alkali to form water as a product.

B Sodium hydroxide will lower the pH of the solution.

C The pH of the neutral solution is 14.

D The pH of the solution before sodium hydroxide is added is below 7.

24. Lead(II) chloride is an insoluble salt.

Which two reagents are used to prepare a pure sample of lead(II) chloride?

A Lead(II) carbonate and dilute hydrochloric acid.

B Lead metal and dilute hydrochloric acid.

C Aqueous lead(II) nitrate and dilute hydrochloric acid.

D Lead(II) oxide and dilute hydrochloric acid.

25. A compound **X**, when heated with an aqueous solution of compound **Y**, produces a gas that turns red litmus blue.

1 **Y** could be potassium hydroxide

2 **X** is an acid

3 **X** could be an ammonium salt

4 **X** could be sodium nitrate

Which statements are correct?

A 1, 2 and 3 **B** 1 and 3 only

C 3 only **D** 2 and 4 only

26. In a neutralisation reaction, which change in particles occurs?

A atoms $\rightarrow$ molecules

B ions $\rightarrow$ molecules

C atoms $\rightarrow$ ions

D ions $\rightarrow$ atoms

37. Which equation represents the reaction between aqueous sodium carbonate and dilute hydrochloric acid?

A $NaCO_3 + HCl \rightarrow NaCl + H_2O + CO_2$

B $NaCO_3 + 2HCl \rightarrow NaCl_2 + H_2O + CO_2$

C $Na_2CO_3 + 2HCl \rightarrow 2NaCl + H_2O + CO_2$

D $Na_2CO_3 + HCl \rightarrow NaCl + H_2O + CO_2$

28. Two solutions are prepared.

- Solution **P** is 0.050 mol/dm^3 hydrochloric acid.
- Solution **Q** is 0.100 mol/dm^3 ethanoic acid.

A 2 cm strip of magnesium ribbon is put into 100 cm^3 of each solution. Effervescence is seen in both solutions but the effervescence is faster in solution **P** than it is in solution **Q**.

Which statement helps to explain this observation?

A Magnesium reacts with solution **P** to form a salt, but does not form a salt with solution **Q**.

B More molecules are dissociated in solution **P** than are dissociated in solution **Q**.

C Solution **Q** contains a stronger acid than solution **P**.

D The particles are closer together in solution **Q** than they are in solution **P**.

29. Which one of the following, when added in **excess** to acidic soil, will increase the pH of the soil to exactly 7.0?

A Calcium carbonate **B** Calcium chloride

C Sodium hydroxide **D** Sodium sulfate

30. Sodium reacts with liquid ammonia at a temperature of –33 °C according to the following balanced chemical equation.

$$2Na(s) + 2NH_3(l) \rightarrow 2NaNH_2(l) + H_2(g)$$

How is the liquid ammonia behaving in this reaction?

A As an acid. **B** As an alkali.

C As an indicator. **D** As a salt.

31. The table below shows information about four different pH indicators.

indicator	pH at which colour change takes place	colour at low pH	colour at high pH
methyl orange	3.2 to 4.4	red	yellow
bromocresol blue	3.8 to 5.4	yellow	blue
methyl red	4.8 to 6.0	red	yellow
bromothymol blue	6.0 to 7.6	yellow	red

Which indicator, or set of indicators, is most precise in showing that a solution has a pH of approximately 4.5 to 4.7?

A Methyl red only

B Bromocresol blue only

C Methyl red and methyl orange

D Methyl red and bromothylmol blue

32. Study the following reactions:

1 hydrochloric acid + copper

2 nitric acid + zinc

3 hydrochloric acid + calcium hydroxide

4 nitric acid + calcium carbonate

Which reactions will produce a gas as one of the products?

A 1 and 2 only **B** 1 and 3 only

C 2 and 3 only **D** 2 and 4 only

33. Acids can be distinguished from alkalis with litmus paper. When a student tested a liquid with red litmus paper, the litmus paper remained red. What correct conclusion can the student make about the liquid?

A It is neither an acid nor an alkali.

B It is not an alkali.

C It must be an acid.

D It must be pure water.

34. What property does hydrochloric acid have?

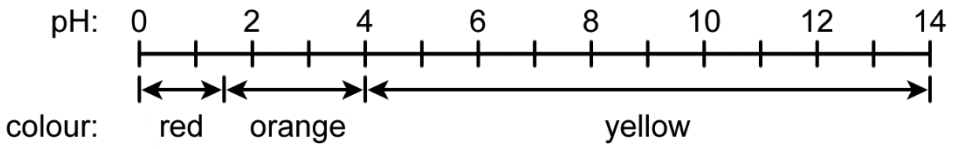
A It liberates ammonia from ammonium salts.

B It reacts with any base to give a salt.

C It reacts with any metal to give hydrogen.

D It turns red litmus paper blue.

35. The colour of indicator **Y** in solutions with different pH values is shown in the diagram below.



Which two solutions can indicator **Y** be used to distinguish between?

A Dilute hydrochloric acid and dilute nitric acid.

B Aqueous sodium hydroxide and aqueous sodium chloride.

C Distilled water and dilute hydrochloric acid.

D Distilled water and aqueous sodium hydroxide.

36. Four aqueous solutions have the pH values shown in the table below.

Solution	**P**	**Q**	**R**	**S**
pH	2	6	8	10

If pairs of the solutions are mixed, which pair **must** produce an acidic mixture?

A **P** and **Q**
B **P** and **R**
C **P** and **S**
D **Q** and **R**

37. Calcium, calcium oxide and calcium carbonate all react with nitric acid at room temperature and pressure. Which of the following statement(s) is true for all three reactions?

1 A salt is produced
2 Water is produced
3 A gas is produced

A 1 only
B 2 only
C 1 and 2 only
D 2 and 3 only

38. Which of the following pairs of substances will react with dilute sulfuric acid to produce a **visible** change?

A Ammonium chloride and sodium chloride.
B Copper and sodium hydroxide.
C Sodium hydroxide and sodium carbonate.
D Zinc and sodium carbonate.

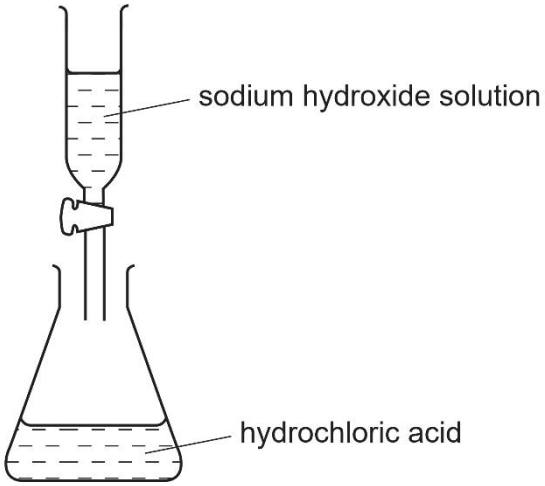
39. Which chemical equation for the reaction between calcium carbonate and phosphoric acid is correct?

A $Ca(CO_3)_2 + H_3PO_4 \rightarrow CaPO_4 + 2H_2O + 2CO_2$
B $Ca_2CO_3 + 2H_3PO_4 \rightarrow 2CaPO_4 + 3H_2O + CO_2$
C $2CaCO_3 + 3H_3PO_4 \rightarrow Ca_3(PO_4)_2 + 2H_2O + 2CO_2$
D $3CaCO_3 + 2H_3PO_4 \rightarrow Ca_3(PO_4)_2 + 3H_2O + 3CO_2$

40. Which reaction will result in the formation of an insoluble salt?

A $Ba(OH)_2 + H_2SO_4 \rightarrow BaSO_4 + 2H_2O$
B $Cu(OH)_2 + H_2SO_4 \rightarrow CuSO_4 + 2H_2O$
C $Fe(OH)_2 + H_2SO_4 \rightarrow FeSO_4 + 2H_2O$
D $2NaOH + H_2SO_4 \rightarrow Na_2SO_4 + 2H_2O$

41. Sodium hydroxide solution was added to dilute hydrochloric acid. The pH of the solution in the flask was measured at intervals until no further change of pH took place.



What would be the pH change of this reaction?

A Decrease to 1

B Decrease to 7

C Increase to 7

D Increase to 12

42. Titration of an acid against a base is a method often used in the preparation of salts.

Which properties of the acid, the base and the salt are required if this method is to be used?

	acid	base	salt
A	insoluble	insoluble	insoluble
B	soluble	insoluble	insoluble
C	soluble	soluble	insoluble
D	soluble	soluble	soluble

43. Which acid and base react together to produce an insoluble salt?

A Hydrochloric acid and sodium hydroxide

B Nitric acid and calcium oxide

C Sulfuric acid and barium hydroxide

D Sulfuric acid and zinc oxide

44. The following statements about dilute sulfuric acid are all correct.

1 Addition of Universal Indicator shows that the solution has a pH value of less than 7.0.

2 A white precipitate is formed when aqueous barium nitrate is added.

3 The solution reacts with copper(II) oxide, forming a blue solution.

4 The solution turns anhydrous copper(II) sulfate from white to blue.

Which two statements confirm the acidic nature of the solution?

A 1 and 2 only **B** 1 and 3 only

C 2 and 4 only **D** 3 and 4 only

45. The table shows the solubility of some compounds of metal **Q** in cold water.

salt	solubility in cold water
carbonate	insoluble
chloride	soluble
sulfate	insoluble

What is metal **Q**?

A Barium **B** Lead

C Magnesium **D** Sodium

46. A metal **M** forms a chloride which dissolves in cold water and has an oxide which dissolves in both strong acids and strong alkalis.

What is **M**?

A Iron **B** Lead

C Sodium **D** Zinc

47. Which statement about oxides is correct?

A A basic oxide is an oxide of a non-metal.

B Acidic oxides contain ionic bonds.

C An amphoteric oxide contains a metal.

D Basic oxides are always gases.

48. How can a pure sample of barium sulfate be obtained from barium carbonate?

A Dissolve it in dilute hydrochloric acid, add dilute sulfuric acid, filter and crystallise.

B Dissolve it in dilute hydrochloric acid, add dilute sulfuric acid, filter and wash.

C Dissolve it in water, add dilute sulfuric acid, filter and crystallise.

D Dissolve it in water, add dilute sulfuric acid, filter and wash.

49. Which row correctly describes the oxides?

	Al_2O_3	K_2O	MgO	SO_2
A	basic	acidic	acidic	amphoteric
B	acidic	basic	amphoteric	acidic
C	amphoteric	basic	amphoteric	acidic
D	amphoteric	basic	basic	acidic

50. Powdered calcium carbonate reacts with dilute hydrochloric acid to produce calcium chloride, water and carbon dioxide.

Which is the correct ionic equation, including state symbols, for this reaction?

A $CaCO_3(s) + 2HCl(aq) \rightarrow CaCl_2(aq) + H_2O(l) + CO_2(g)$

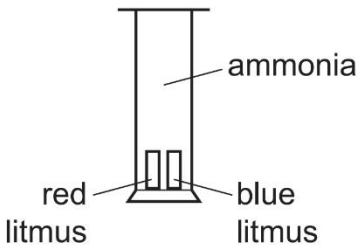
B $Ca^{2+}(aq) + CO_3^{2-}(aq) + 2H^+(aq) \rightarrow Ca^{2+}(aq) + H_2O(l) + CO_2(g)$

C $CO_3^{2-}(aq) + 2H^+(aq) \rightarrow H_2O(l) + CO_2(g)$

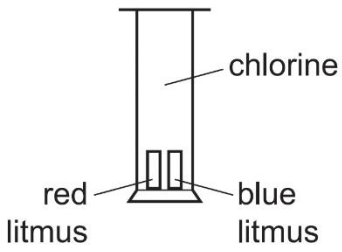
D $CaCO_3(s) + 2H^+(aq) \rightarrow Ca^{2+}(aq) + H_2O(l) + CO_2(g)$

51. Four gas jars each contain one of the gases ammonia, chlorine, hydrogen chloride and oxygen. A strip of damp blue litmus paper and a strip of damp red litmus paper are placed in each jar. In which gas jar will both the damp blue litmus paper and the damp red litmus paper change colour?

A



B



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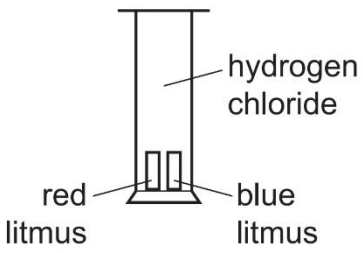
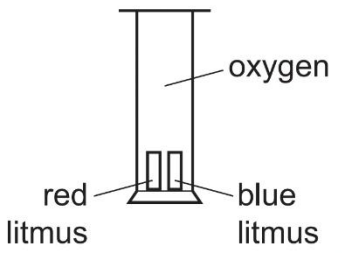


D



52. Which uses for sulfuric acid are correct.

1 As a bleach in the manufacture of wood pulp for paper.

2 As a food preservative in tinned foods.

3 As a raw material in the manufacture of detergents.

4 As a fertiliser.

A 1 and 3 only **B** 2 and 4 only

C 2 only **D** 3 only

53. Which ion is **not** present in dilute sulfuric acid?

A H^+ **B** OH^- **C** SO_3^{2-} **D** SO_4^{2-}

54. Which statement about weak acids is correct?

A They are partially ionised.

B They do not react with metals.

C They do not react with strong alkalis.

D Their solutions have pH values in the range 0 to 2.

55. A household cleaning compound is used to remove calcium carbonate from bathroom surfaces. The compound reacts with the calcium carbonate to form a soluble salt, carbon dioxide and water. What is the pH of this cleaning compound?

A pH 2 **B** pH 7 **C** pH 10 **D** pH 14

56. Dilute hydrochloric acid is added separately to samples of copper, copper(II) oxide and copper(II) carbonate. Which row correctly shows whether copper(II) chloride is produced?

	Cu	CuO	$CuCO_3$
A	✓	✓	✓
B	✗	✓	✗
C	✓	✗	✓
D	✗	✓	✓

key:

✓ = copper(II) chloride produced

✗ = copper(II) chloride not produced

57. Which ions are present when hydrochloric acid has exactly neutralised aqueous sodium hydroxide?

A Na^+, Cl^-, H^+ and OH^-

B Na^+, Cl^- and H^+ only

C Na^+ and Cl^- only

D H^+ and OH^- only

58. Which experiment will result in the formation of a white precipitate?

A Aqueous barium nitrate added to aqueous sodium chloride.

B Aqueous sodium carbonate added to aqueous calcium chloride.

C Carbon dioxide passed through aqueous potassium chloride.

D Dilute hydrochloric acid added to aqueous ammonia.

59. How should a pure, dry sample of barium sulfate be prepared from a sample of barium carbonate?

A Titrate sulfuric acid against the barium carbonate using a suitable indicator. Heat the resulting solution until it is saturated and then crystallise the barium sulfate from solution.

B Add excess sulfuric acid to the barium carbonate. Filter the mixture. Wash and dry the residue.

C Add excess aqueous sodium sulfate to the barium carbonate. Filter the mixture. Wash and dry the residue.

D Add excess nitric acid to the barium carbonate. Add excess sodium sulfate to the resulting solution. Filter the mixture. Wash and dry the residue.

60. Why is ethanoic acid described as a weak acid?

A It is an organic acid.

B It is a poor conductor of electricity.

C It is only slightly dissociated in water.

D It reacts only with very reactive metals.

61. What is the best method to prepare a sample of silver chloride?

A Add silver nitrate to chlorine.

B Add silver to hydrochloric acid.

C Burn silver in chlorine.

D Mix aqueous solutions of silver nitrate and sodium chloride.

62. Solution **T** has the following properties.

1 It reacts with magnesium forming a gas.

2 It reacts with calcium carbonate forming a gas.

Which statement about solution **T** is correct?

A It contains more OH^- ions than H^+ ions.

B It has pH 9.

C Its reaction with calcium carbonate produces hydrogen.

D It reacts with aqueous ammonia.

63. Which row correctly shows whether the hydrogen ion concentration and the pH of ethanoic acid are higher or lower than those of hydrochloric acid of the same concentration?

	hydrogen ion concentration	pH
A	higher	higher
B	higher	lower
C	lower	higher
D	lower	lower

64. Which salts could be prepared by precipitation?

1 barium sulfate

2 lead(II) chloride

3 copper(II) chloride

4 zinc sulfate

A 1 and 2 only

B 3 and 4 only

C 1 and 3 only

D 2 and 4 only

65. Silver is below hydrogen in the reactivity series.

Which row describes the chemicals used and the method of separation used to prepare a pure sample of silver chloride?

	chemicals used	method of separation of product
A	silver and hydrochloric acid	crystallisation
B	silver and hydrochloric acid	filtration
C	silver nitrate and hydrochloric acid	crystallisation
D	silver nitrate and hydrochloric acid	filtration

66. The water in a lake is acidic and the fish are dying. The water in the lake needs to be neutralised. Which compound can be added in excess to neutralise the water in the lake?

A Calcium carbonate

B Phosphoric acid

C Potassium hydroxide

D Sodium nitrate

67. Two incomplete statements about the preparation of an insoluble salt are given.

.....1..... can be used to prepare insoluble salts, such as .....2..... .

The salt is collected by ..... 3..... and it is then .....4..... .

Which words correctly complete gaps 1 - 4?

	1	2	3	4
A	precipitation	barium nitrate	filtration	evaporated
B	precipitation	lead(II) sulfate	evaporation	washed and dried
C	precipitation	lead(II) sulfate	filtration	washed and dried
D	titration	barium nitrate	evaporation	washed and dried

68. Which statement about acids and bases is correct?

A All strong acids react with carbonates, but all weak acids do not.

B The oxides of Group 1 metals are amphoteric.

C The pH of 1.0 mol/dm^3 ethanoic acid, CH_3COOH, is higher than the pH of 1.0 mol/dm^3 sulfuric acid, H_2SO_4.

D The pH of 1.0 mol/dm^3 nitric acid, HNO_3, is lower than the pH of 1.0 mol/dm^3 hydrochloric acid, HC*l*.

69. Which method should be used to make a pure sample of potassium chloride?

A Adding AgC*l*(s) to KNO_3(aq)

B Adding excess K_2CO_3(s) to HC*l*(aq)

C Mixing KNO_3(aq) with NaC*l*(aq)

D Titrating KOH(aq) with HC*l*(aq)

70. A pure sample of lead(II) sulfate is made by reacting aqueous solutions of two salts. The lead(II) sulfate formed is then separated from the mixture.

Which solutions and method of separation are used?

	salt solution 1	salt solution 2	salt solution 3
A	lead(II) chloride	sodium sulfate	crystallisation
B	lead(II) chloride	sodium sulfate	filtration
C	lead(II) nitrate	potassium sulfate	crystallisation
D	lead(II) nitrate	potassium sulfate	filtration

71. Which row shows the correct state symbols for the reaction between calcium carbonate and dilute hydrochloric acid? (The conditions are room temperature and pressure.)

	$CaCO_3$ + HCl → $CaCl_2$ + H_2O + CO_2				
A	(s)	(aq)	(s)	(aq)	(g)
B	(s)	(*l*)	(aq)	(*l*)	(g)
C	(s)	(*l*)	(s)	(aq)	(g)
D	(s)	(aq)	(aq)	(*l*)	(g)

72. Aqueous calcium hydroxide is an alkali. It is neutralised by dilute nitric acid to produce calcium nitrate and water.

What is the **ionic** equation for this reaction?

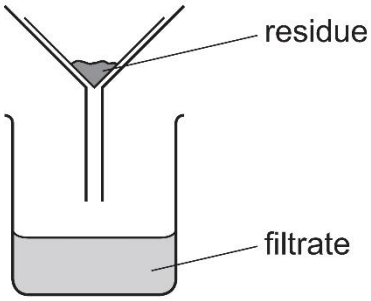
A $Ca^+ + OH^- + H^+ + NO_3^- \rightarrow CaNO_3 + H_2O$

B $Ca(OH)_2 + 2HNO_3 \rightarrow Ca(NO_3)_2 + 2H_2O$

C $Ca^{2+}(OH^-)_2 + 2H^+NO_3^- \rightarrow Ca^{2+}(NO_3^-)_2 + 2H_2O$

D $OH^- + H^+ \rightarrow H_2O$

73. Pure lead(II) sulfate is prepared by mixing two substances, **X** and **Y**. When the reaction is complete the mixture is filtered. Pure lead(II) sulfate is obtained.



Which row shows the best way to prepare lead(II) sulfate?

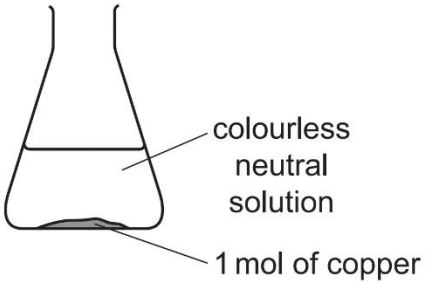
	substance **X**	substance **Y**	method after filtration
A	aqueous lead(II) nitrate	aqueous sodium sulfate	crystallise the filtrate
B	aqueous lead(II) nitrate	aqueous sodium sulfate	wash and dry the residue
C	solid lead(II) carbonate	dilute sulfuric acid	crystallise the filtrate
D	solid lead(II) carbonate	dilute sulfuric acid	wash and dry the residue

74. Which method of preparation would be suitable for making these salts?

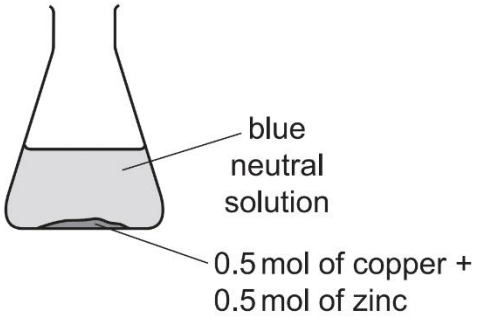
	titration	metal + acid	metal carbonate + acid	precipitation
A	sodium nitrate	lead(II) chloride	copper(II) sulfate	zinc sulfate
B	zinc sulfate	copper(II) sulfate	sodium nitrate	lead(II) chloride
C	zinc sulfate	sodium nitrate	lead(II) chloride	copper(II) sulfate
D	sodium nitrate	zinc sulfate	copper(II) sulfate	lead(II) chloride

75. In an experiment, 1 mol of powdered copper and 1 mol of powdered zinc are placed in a flask.
Dilute acid, containing 1 mol of acid, is added to the flask.
The flask is left until all the reactions, if any, are complete.
Which diagram shows the result of the experiment?

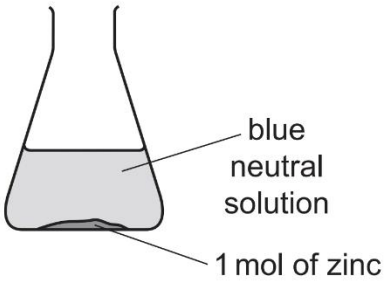
A



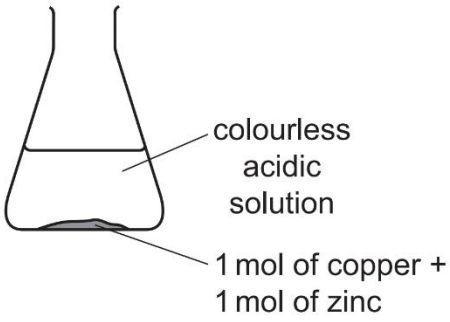
B



C



D



- Scan the QR Code below to view the answers to this assignment.

http://www.chemist.sg/acids/acids_assignments/acids_assignment_6_ans.pdf