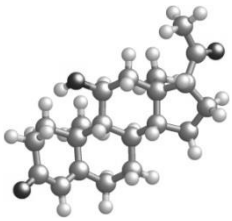

Chem!stry

Name: …… **( )**

Class: ……

Date: …… **/** …… **/** ……

Electrochemistry – True or False?

1.	During the electrolysis of molten $PbBr_{2(l)}$, lead(II) ions are oxidised at the anode:	True ☐	False ☐
2.	During electrolysis, reduction takes place at the cathode:	True ☐	False ☐
3.	During the electrolysis of aqueous $Ca(NO_3)_{2(aq)}$, hydrogen ions are selectively reduced at the cathode:	True ☐	False ☐
4.	During electrolysis, mobile electrons carry charge through the solution:	True ☐	False ☐
5.	During the electrolysis of aqueous $CuSO_{4(aq)}$ using copper electrodes, the anode decreases in mass:	True ☐	False ☐
6.	During the electrolysis of dilute aqueous $KCl_{(aq)}$, chlorine ions are selectively oxidised at the anode:	True ☐	False ☐
7.	During the electrolysis of aqueous $AgNO_{3(aq)}$ using graphite electrodes, the anode decreases in mass:	True ☐	False ☐
8.	During the electrolysis of concentrated aqueous $NaBr_{(aq)}$, bromide ions are selectively oxidised at the anode:	True ☐	False ☐
9.	When a zinc half-cell is connected to a copper half-cell, electrons flow from the zinc towards the copper:	True ☐	False ☐
10.	When a zinc half-cell is connected to a silver half-cell, mobile ions carry charge through the salt bridge:	True ☐	False ☐
11.	When a zinc half-cell is connected to a copper half-cell, oxidation takes place in the copper half-cell:	True ☐	False ☐
12.	When a magnesium half-cell is connected to a zinc half-cell, reduction takes place in the zinc half-cell:	True ☐	False ☐
13.	When a zinc half-cell is connected to a copper half-cell, the potential difference produced is greater than when a zinc half-cell is connected to a silver half-cell:	True ☐	False ☐
14.	When a zinc half-cell is connected to a copper half-cell, the electrons flow in the same direction as they would when a zinc half-cell is connected to a magnesium half-cell:	True ☐	False ☐
15.	When a zinc half-cell is connected to a copper half-cell, oxidation takes place at the anode:	True ☐	False ☐
16.	When a zinc half-cell is connected to a copper half-cell, the potential difference eventually drops to zero once all of the copper(II) ions have been reduced:	True ☐	False ☐

- Scan the QR code given below to view the answers to this assignment.

http://www.chemist.sg/electro_chem/electrochem_true_false_ans.pdf