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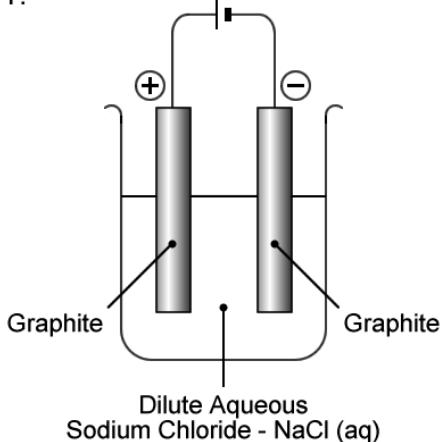
Class:

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Revision – Electrolysis of Aqueous Salts and Acids

- For the electrolysis of each aqueous salt; **a)** write the ionic half-equation for the reaction at the anode, **b)** write the ionic half-equation for the reaction at the cathode, **c)** state the name and formula of the compound that remains in aqueous solution.

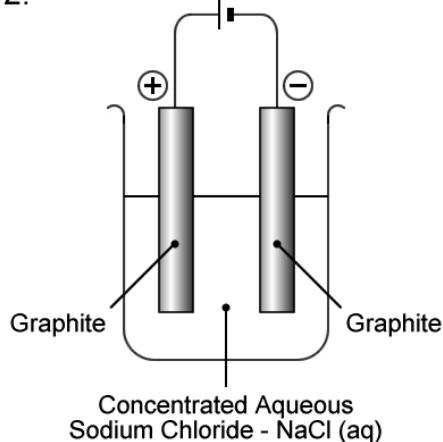
1.



Dilute Aqueous Sodium Chloride - NaCl (aq)

- a)** Anode (+)
- b)** Cathode (-)
- c)** In Solution

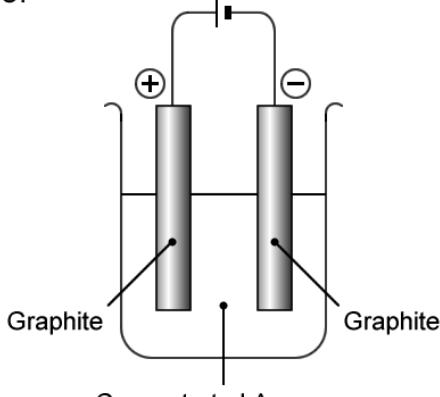
2.



Concentrated Aqueous Sodium Chloride - NaCl (aq)

- a)** Anode (+)
- b)** Cathode (-)
- c)** In Solution

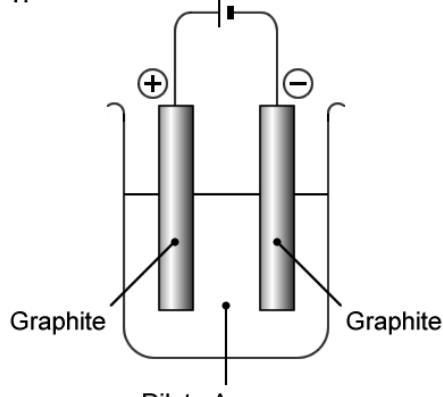
3.



Concentrated Aqueous Potassium Iodide - KI (aq)

- a)** Anode (+)
- b)** Cathode (-)
- c)** In Solution

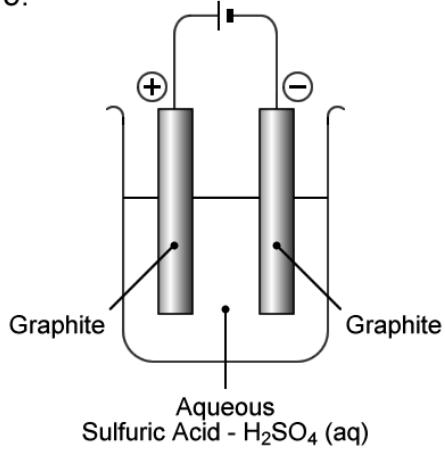
4.



Dilute Aqueous Potassium Iodide - KI (aq)

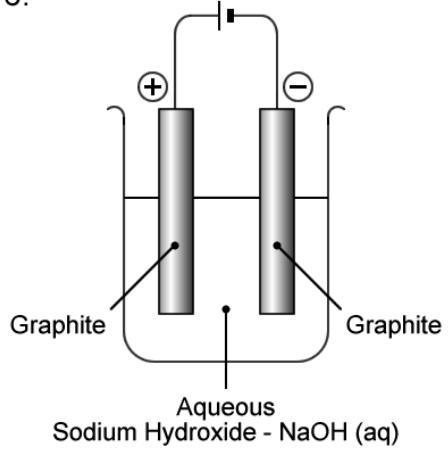
- a)** Anode (+)
- b)** Cathode (-)
- c)** In Solution

5.



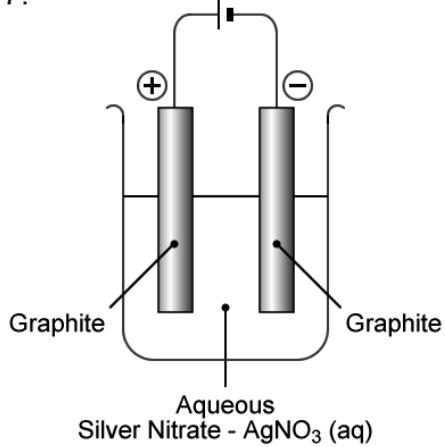
- a) Anode (+)
- b) Cathode (-)
- c) In Solution

6.



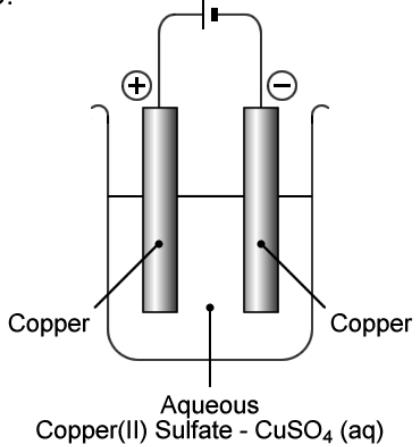
- a) Anode (+)
- b) Cathode (-)
- c) In Solution

7.



- a) Anode (+)
- b) Cathode (-)
- c) In Solution

8.



- a) Anode (+)
- b) Cathode (-)
- c) In Solution

- Scan the QR code given below to view the answers to this assignment.



http://www.chemist.sg/electro_chem/electrolysis_aqueous_salts_ans.pdf