

Name: ()

Chem!stry Class:

Date: / /

Revision – Electrolysis of Molten Binary Salts

- For the electrolysis of each molten salt; **a)** write the ionic half-equation for the reaction at the anode, **b)** state what you would observe at the anode, **c)** write the ionic half-equation for the reaction at the cathode, **d)** state what you would observe at the cathode.

1.

Graphite (+) Graphite (-)
Molten Calcium Oxide - CaO (l)

a) Anode (+)

b) Observation

c) Cathode (-)

d) Observation

2.

Graphite (+) Graphite (-)
Molten Zinc Chloride - ZnCl₂ (l)

a) Anode (+)

b) Observation

c) Cathode (-)

d) Observation

3.

Graphite (+) Graphite (-)
Molten Tin(II) Chloride - SnCl₂ (l)

a) Anode (+)

b) Observation

c) Cathode (-)

d) Observation

4.

Graphite (+) Graphite (-)
Molten Caesium Bromide - CsBr (l)

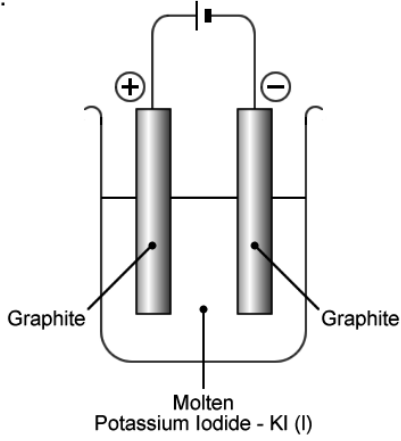
a) Anode (+)

b) Observation

c) Cathode (-)

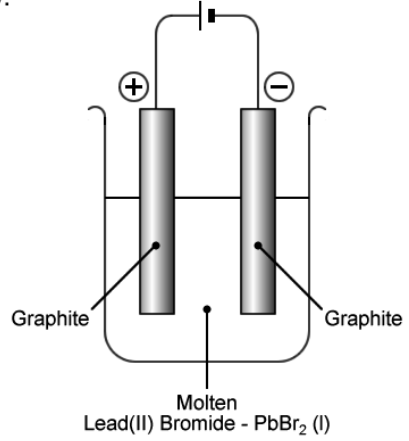
d) Observation

5.



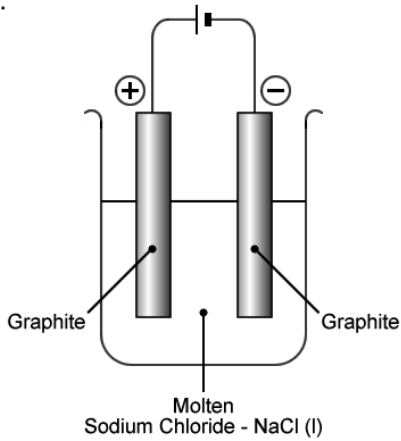
- a) Anode (+)
- b) Observation
- c) Cathode (-)
- d) Observation

6.



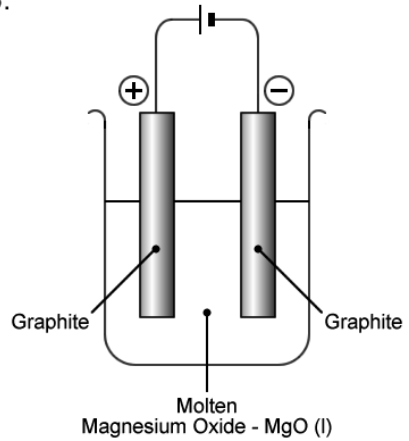
- a) Anode (+)
- b) Observation
- c) Cathode (-)
- d) Observation

7.



- a) Anode (+)
- b) Observation
- c) Cathode (-)
- d) Observation

8.



- a) Anode (+)
- b) Observation
- c) Cathode (-)
- d) Observation

- Scan the QR code given below to view the answers to this assignment.



http://www.chemist.sg/electro_chem/electrolysis_molten_salts_ans.pdf