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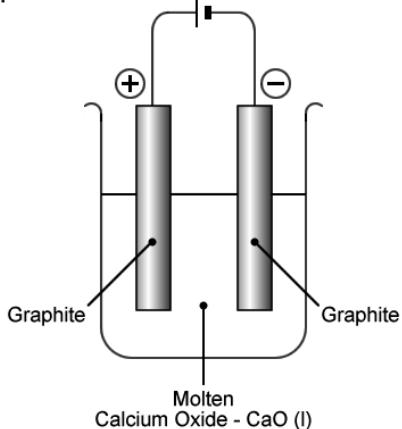
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Revision – Electrolysis of Molten Binary Salts

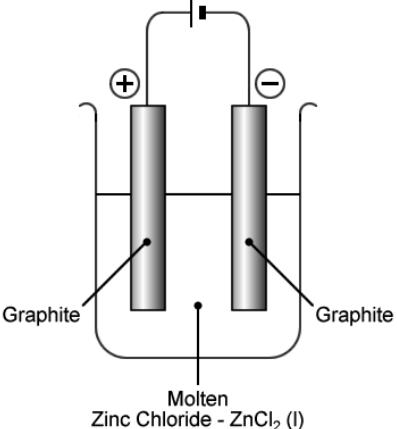
- For the electrolysis of each molten salt; **a)** write the ionic half-equation for the reaction at the anode, **b)** state what you would observe at the anode, **c)** write the ionic half-equation for the reaction at the cathode, **d)** state what you would observe at the cathode.

1.

Molten Calcium Oxide - $\text{CaO} (\text{l})$

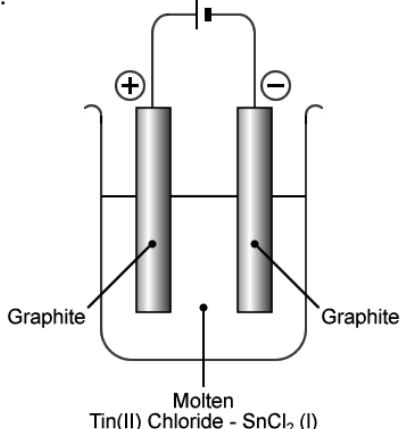
- a)** Anode (+)
b) Observation
c) Cathode (-)
d) Observation

2.

Molten Zinc Chloride - $\text{ZnCl}_2 (\text{l})$

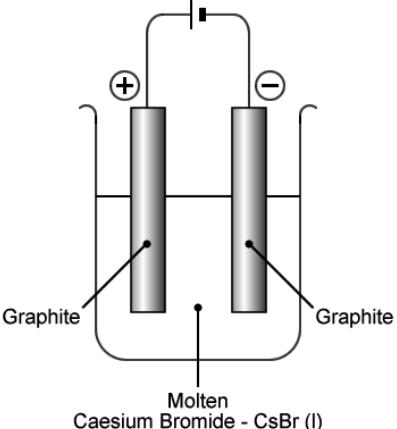
- a)** Anode (+)
b) Observation
c) Cathode (-)
d) Observation

3.

Molten Tin(II) Chloride - $\text{SnCl}_2 (\text{l})$

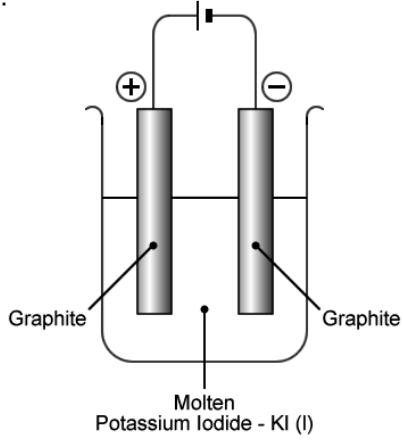
- a)** Anode (+)
b) Observation
c) Cathode (-)
d) Observation

4.

Molten Caesium Bromide - $\text{CsBr} (\text{l})$

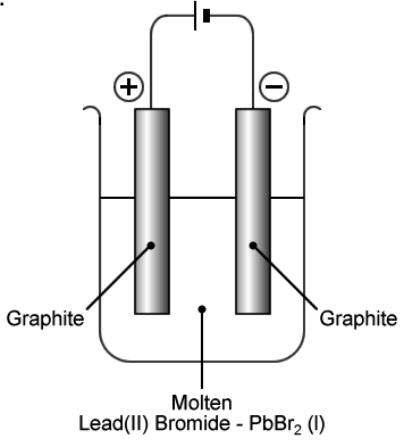
- a)** Anode (+)
b) Observation
c) Cathode (-)
d) Observation

5.



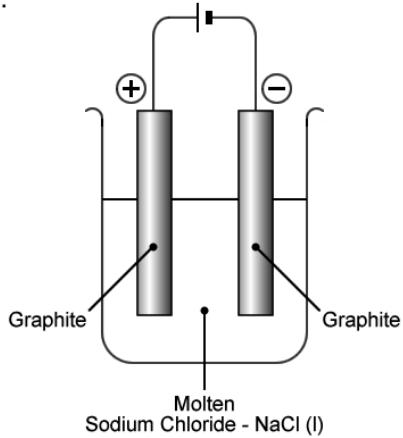
- a) Anode (+)
- b) Observation
- c) Cathode (-)
- d) Observation

6.



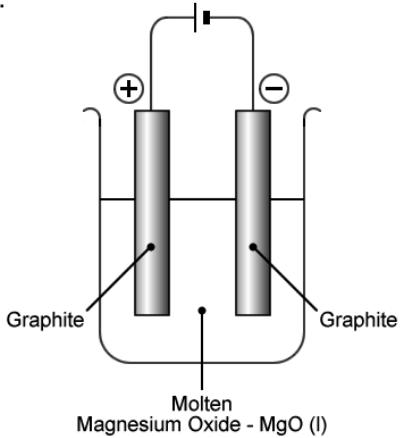
- a) Anode (+)
- b) Observation
- c) Cathode (-)
- d) Observation

7.



- a) Anode (+)
- b) Observation
- c) Cathode (-)
- d) Observation

8.



- a) Anode (+)
- b) Observation
- c) Cathode (-)
- d) Observation

- Scan the QR code given below to view the answers to this assignment.



http://www.chemist.sg/electro_chem/electrolysis_molten_salts_ans.pdf