



Name: ..... ( )

*Chem!stry* Class: .....

Date: ..... / ..... / .....

### Revision – Simple Cells (Batteries)

- For each one of the following simple cells; **a)** label the two boxes to correctly identify the anode and the cathode, **b)** write the ionic half-equation for the reaction at the anode, **c)** write the ionic half-equation for the reaction at the cathode, **d)** draw an arrow over the voltmeter to clearly indicate the direction in which electrons are flowing through the external circuit.

1.

Iron                                  Copper

Aqueous Copper(II) Sulfate -  $\text{CuSO}_{4(aq)}$

**b) Anode** .....

**c) Cathode** .....

2.

Copper                                  Zinc

Aqueous Copper(II) Sulfate -  $\text{CuSO}_{4(aq)}$

**b) Anode** .....

**c) Cathode** .....

3.

Silver                                  Iron

Aqueous Silver Nitrate -  $\text{AgNO}_{3(aq)}$

**b) Anode** .....

**c) Cathode** .....

4.

Zinc                                  Lead

Aqueous Lead(II) Nitrate -  $\text{Pb(NO}_{3})_{2(aq)}$

**b) Anode** .....

**c) Cathode** .....

- Scan the QR code given below to view the answers to this assignment.



[http://www.chemist.sg/electro\\_chem/simple\\_cells\\_ans.pdf](http://www.chemist.sg/electro_chem/simple_cells_ans.pdf)