



# Chem!stry

Name: ..... ( )

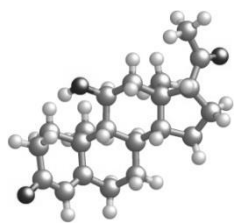
Class: .....

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## Chemical Formulae

- Using your valency table, complete the formulae of the compounds that are formed through combinations of the following cations (positively charged ions) and anions (negatively charged ions). **Note:** The charge on an ion represents the valency of the ion and should *not* be included when writing the final formula of the compound, e.g. sodium chloride is written as  $\text{NaCl}$  and *not* as  $\text{Na}^+\text{Cl}^-$ .

		Anions							
Cations	$\text{Cl}^-$	$\text{Br}^-$	$\text{I}^-$	$\text{OH}^-$	$\text{NO}_3^-$	$\text{O}^{2-}$	$\text{CO}_3^{2-}$	$\text{SO}_4^{2-}$	$\text{PO}_4^{3-}$
$\text{Na}^+$									
$\text{K}^+$									
$\text{NH}_4^+$									
$\text{Ag}^+$									
$\text{Ca}^{2+}$									
$\text{Mg}^{2+}$									
$\text{Fe}^{2+}$									
$\text{Cu}^{2+}$									
$\text{Pb}^{2+}$									
$\text{Fe}^{3+}$									
$\text{Al}^{3+}$									



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## Writing Chemical Formulae

• Use the valency table to write the chemical formulae of the following compounds:

- a) sodium chloride .....
- b) sodium oxide .....
- c) potassium bromide .....
- d) magnesium chloride .....
- e) magnesium oxide .....
- f) aluminium chloride .....
- g) aluminium oxide .....
- h) copper(II) oxide .....
- i) iron(III) chloride .....
- j) potassium nitrate .....
- k) calcium nitrate .....
- l) sodium carbonate .....
- m) magnesium carbonate .....
- n) copper(II) sulfate .....
- o) copper(II) nitrate .....
- p) silver nitrate .....
- q) lead(II) bromide .....
- r) lead(II) nitrate .....
- s) ammonium chloride .....
- t) ammonium nitrate .....
- u) ammonium sulfate .....
- v) silver oxide .....
- w) zinc sulfate .....
- x) zinc nitrate .....
- y) nitric acid .....
- z) sulfuric acid .....

- Scan the QR code below for the answers to this assignment.



[http://www.chemist.sg/formulae\\_equations/formulae\\_ans.pdf](http://www.chemist.sg/formulae_equations/formulae_ans.pdf)