

Chem!stry

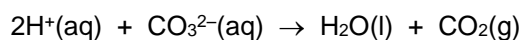
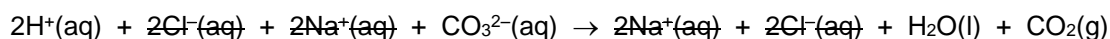
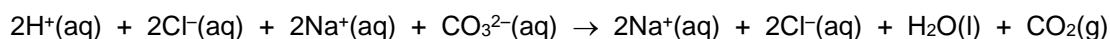
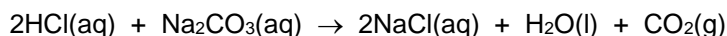
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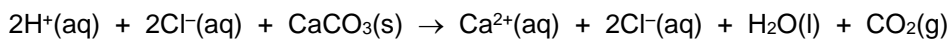
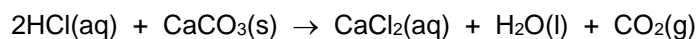
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Ionic Equations

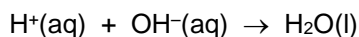
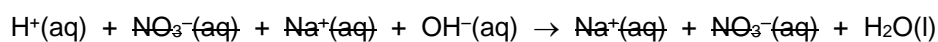
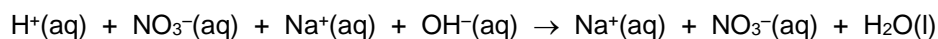
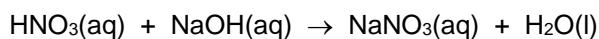
Example One:



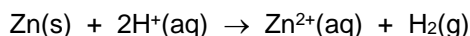
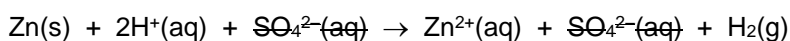
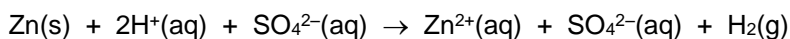
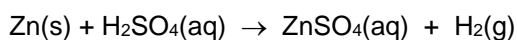
Example Two:



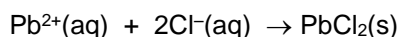
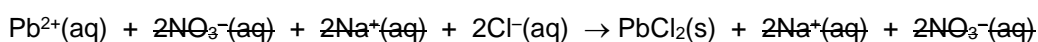
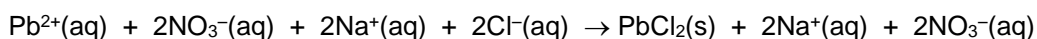
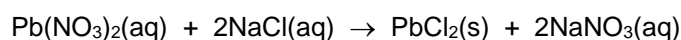
Example Three:



Example Four:

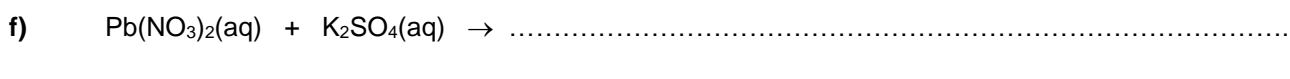
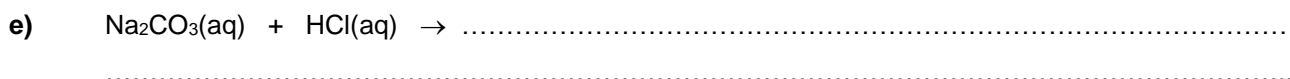
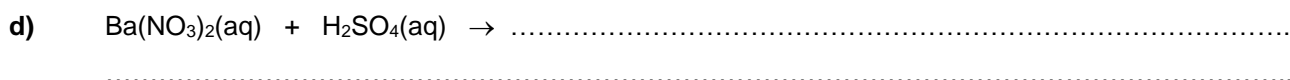
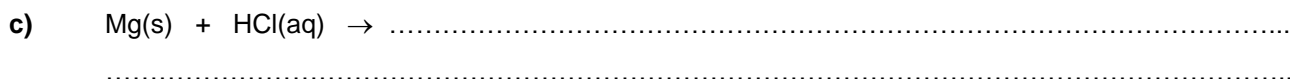
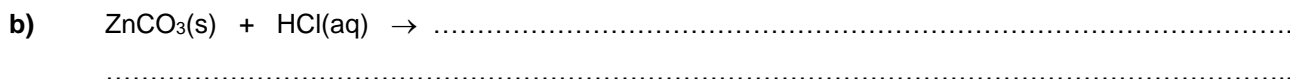
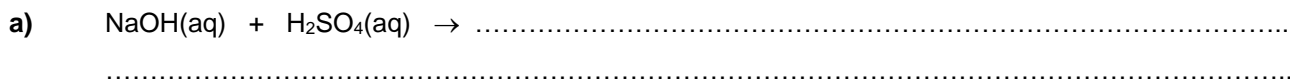


Example Five:



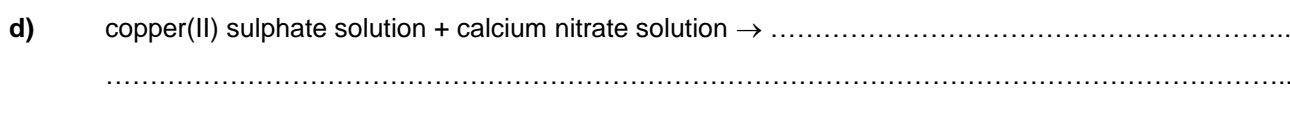
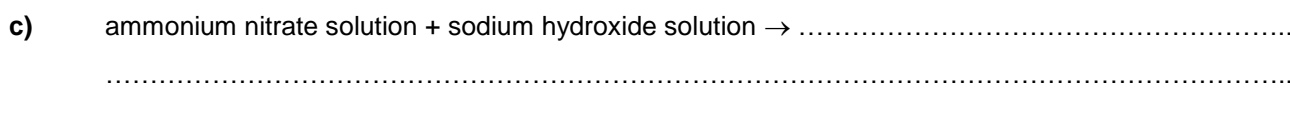
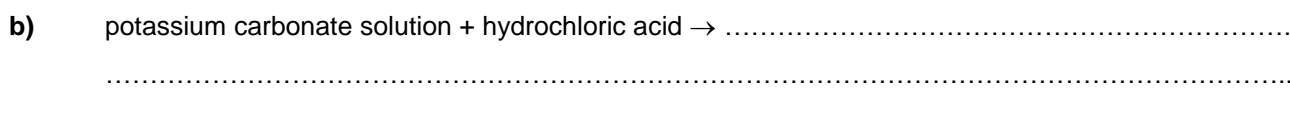
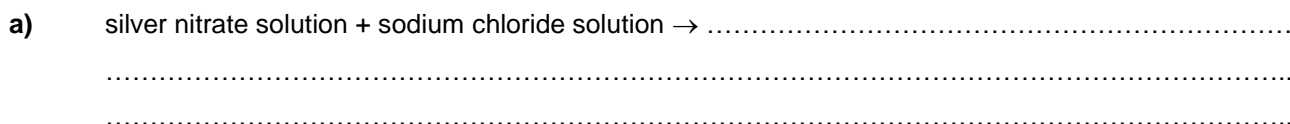
Question One:

Complete and balance each one of the following chemical equations and then write the ionic equation for the reaction:



Question Two:

For each one of the following word equations, write both the balanced chemical equation as well as the ionic equation:



e) magnesium + sulphuric acid →
.....
.....

f) zinc oxide + nitric acid →
.....
.....

g) sodium hydroxide solution + iron(II) nitrate solution →
.....
.....

h) lead(II) nitrate solution + sodium sulphate solution →
.....
.....

i) barium nitrate solution + magnesium sulphate solution →
.....
.....

j) sodium hydroxide solution + phosphoric acid →
.....
.....

k) ammonium chloride solution + sodium hydroxide →
.....
.....

l) copper(II) sulfate solution + sodium hydroxide →
.....
.....

- Scan the QR code below for the answers to this assignment.



http://www.chemist.sg/formulae_equations/ionic_equ_ans.pdf