



Chem!stry

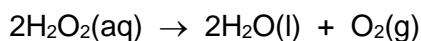
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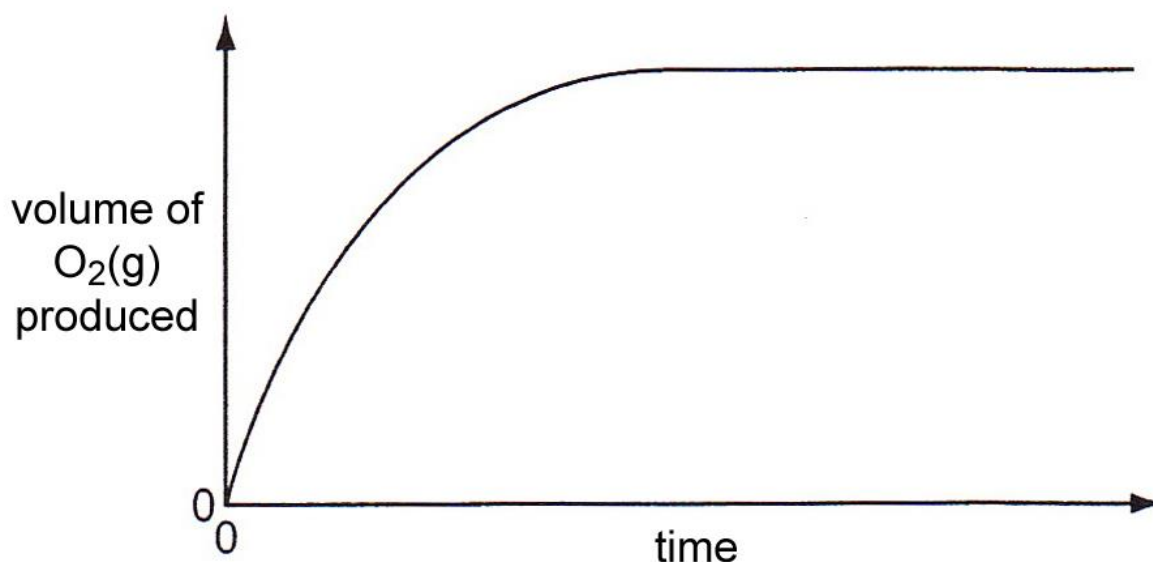
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Rate of Reaction Graphs – Two

- Hydrogen peroxide decomposes into water and oxygen according to the following balanced chemical equation:



- The graph given below is for the decomposition of 20.0 cm³ of 1.0 mol/dm³ hydrogen peroxide at room temperature, without a catalyst.



- On the axes given above, sketch the graph that would be obtained for the decomposition of 20.0 cm³ of 1.0 mol/dm³ hydrogen peroxide at 50.0 °C, without a catalyst. Label **A**.
- On the axes given above, sketch the graph that would be obtained for the decomposition of 20.0 cm³ of 1.0 mol/dm³ hydrogen peroxide at 10.0 °C, without a catalyst. Label **B**.
- On the axes given above, sketch the graph that would be obtained for the decomposition of 20.0 cm³ of 1.0 mol/dm³ hydrogen peroxide at room temperature, with a catalyst. Label **C**.
- On the axes given above, sketch the graph that would be obtained for the decomposition of 10.0 cm³ of 1.0 mol/dm³ hydrogen peroxide at room temperature, without a catalyst. Label **D**.
- On the axes given above, sketch the graph that would be obtained for the decomposition of 20.0 cm³ of 0.5 mol/dm³ hydrogen peroxide at room temperature, without a catalyst. Label **E**.
- On the axes given above, sketch the graph that would be obtained for the decomposition of 7.50 cm³ of 2.0 mol/dm³ hydrogen peroxide at room temperature, without a catalyst. Label **F**.