

# Chem!stry

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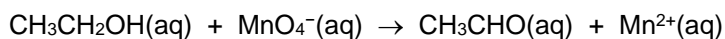
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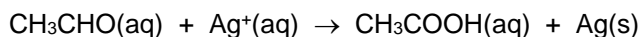
## Ionic Equations for Redox Reactions – Advanced

- For each one of the following reactions, deduce two logical ionic half-equations (one for each reactant). Next, combine the two ionic half-equations together in order to produce a single ionic equation that accurately describes the chemical changes that are taking place.
- For each reaction, identify the species that has been oxidised and the species that has been reduced.
- For each reaction, identify the oxidising agent and the reducing agent.

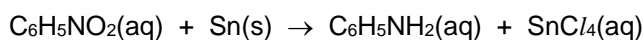
### Question One:



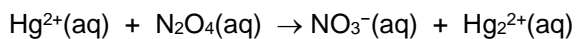
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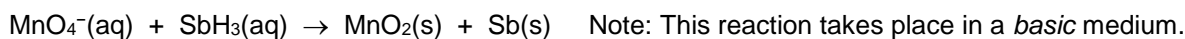
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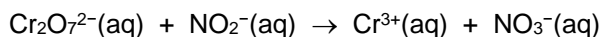
### Question Four:



### Question Five:



### Question Six:



- Scan the QR code below for the answers to this assignment.



[http://www.chemist.sg/redox/advanced\\_ionic\\_equations\\_ans.pdf](http://www.chemist.sg/redox/advanced_ionic_equations_ans.pdf)